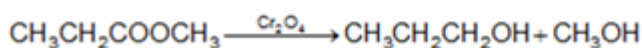
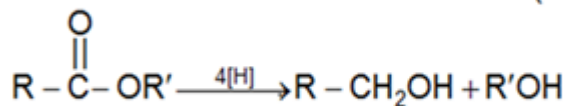
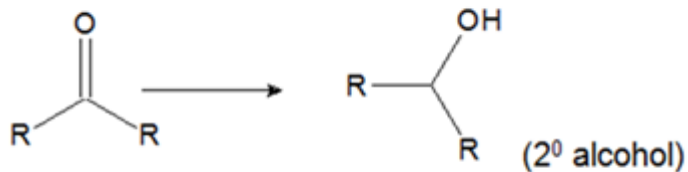
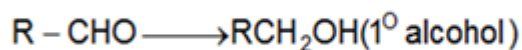
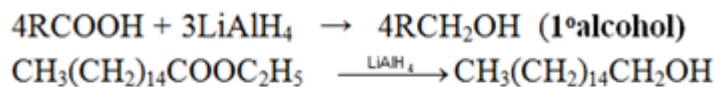


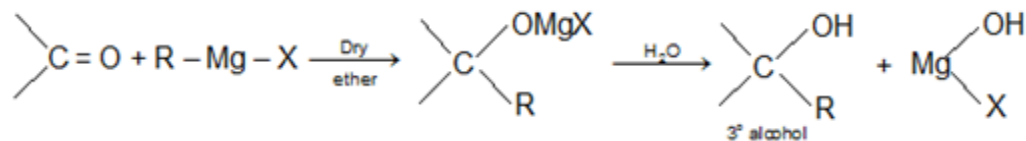
## Alcohols

### Preparation of Alcohols:

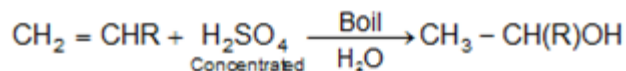
- **By hydrolysis of haloalkanes** :  $R-X + \text{aq. KOH} \rightarrow \text{ROH} + \text{KX}$
- **By reduction of Carbonyl compounds**



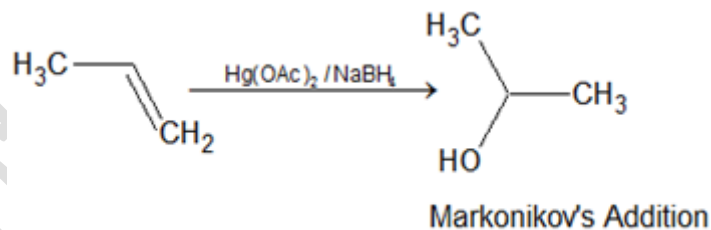
- **By the action of Grignard's Reagent on aldehydes, ketones and esters**



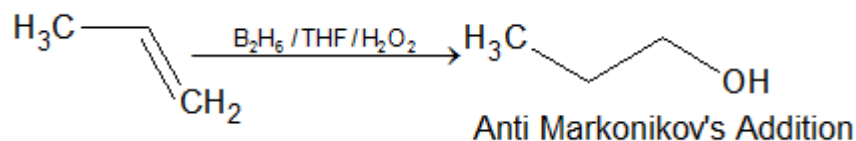
- **By Aliphatic Primary Amines**:  $\text{RCH}_2\text{NH}_2 + \text{HNO}_2 \rightarrow \text{RCH}_2\text{OH} + \text{N}_2 + \text{H}_2\text{O}$
- **Hydration of alkenes:**



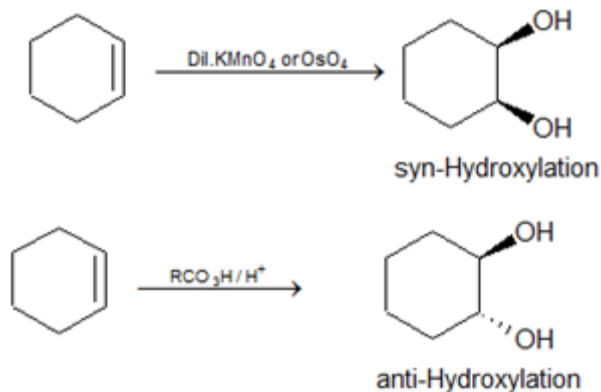
- **Oxymercuration-demercuration:**



- **Hydroboration-oxidation:**



- **Hydroxylation of alkenes:**



### Physical Properties of Alcohol:

Lower alcohols are liquid at room temperature while higher ones are solid.

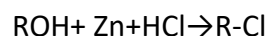
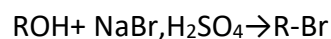
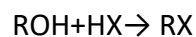
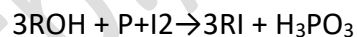
High boiling point due to presence of intermolecular hydrogen bonding. Order of Boiling Point: primary > secondary > tertiary

Solubility in water decreases with increase in molecular mass due to decrease in extent of intermolecular hydrogen bonding.

### Chemical Properties of Alcohol:

**Alcohol's reaction with metal:**  $\text{ROH} + \text{Na} \rightarrow 2\text{RO}^+\text{Na}^- + \text{H}_2$

**Formation of Halides:**



**Reaction with  $\text{HNO}_3$ :**  $\text{R-OH} + \text{HO-NO}_2 \rightarrow \text{R-O-NO}_2$

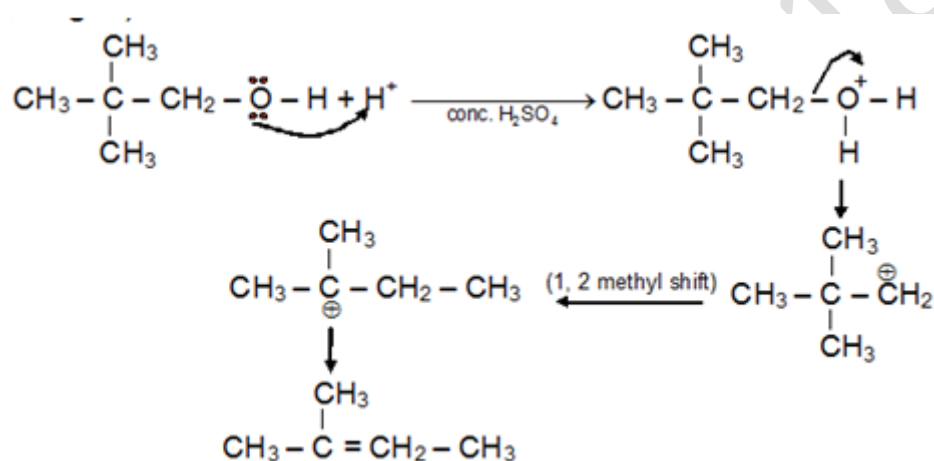
**Reaction with carboxylic acid (Esterification) :**  $R-OH + R'-COOH + H^+ \leftrightarrow R'-COOR$

**Reaction with Grignard reagent:**  $R'OH + RMgX \rightarrow RH + R'OMgX$

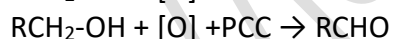
**Reduction of alcohol :**  $ROH + 2HI + Red P \rightarrow RH + I_2 + H_2O$

**Dehydration of Alcohol:** Dehydration of alcohols takes place in acidic medium. Intra-molecular dehydration leads to the formation of alkene while inter molecular dehydration which forms ether. Ease of dehydration:  $3^\circ > 2^\circ > 1^\circ$

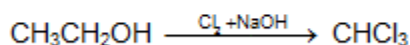
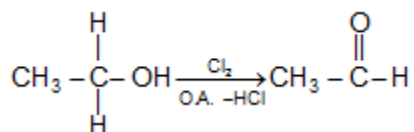
**Satzeff's Rule :** Elimination through  $\beta$  carbon containing minimum  $\beta$  hydrogen



**Oxidation of Alcohol:**



**Haloform Reaction:** Compound containing  $\text{CH}_3\text{CO}-$  group (or compound on oxidation gives  $\text{CH}_3\text{CO}-$  group) which is attached with a C or H, in presence of halogen and mild alkali gives haloform.  $\text{CH}_3\text{-CH}_2\text{-COCH}_2\text{-CH}_3$ ,  $\text{CH}_3\text{-CO-Cl}$ ,  $\text{CH}_3\text{COOH}$  will not respond to haloform reaction while  $\text{CH}_3\text{CH}_2\text{OH}$  will respond to haloform Reaction.



**Test for Alcohols:**

1. Lucas Test:

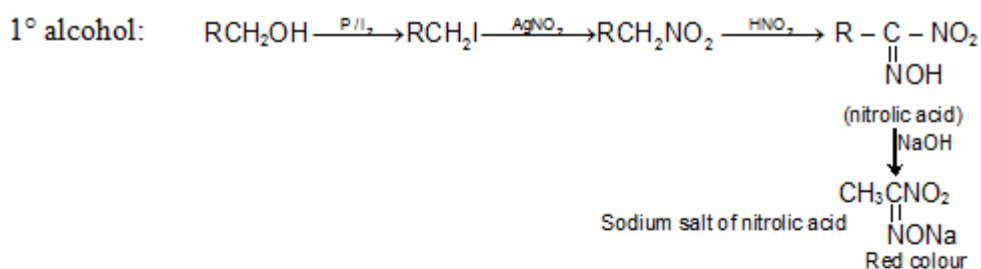
Alcohols + ZnCl<sub>2</sub> + HCl

1° Alcohol: RCH<sub>2</sub>OH + ZnCl<sub>2</sub> + HCl → No reaction at room temperature

2° Alcohol: R<sub>2</sub>CHOH + ZnCl<sub>2</sub> + HCl → R<sub>2</sub>CHCl White turbidity after 5-10 min.

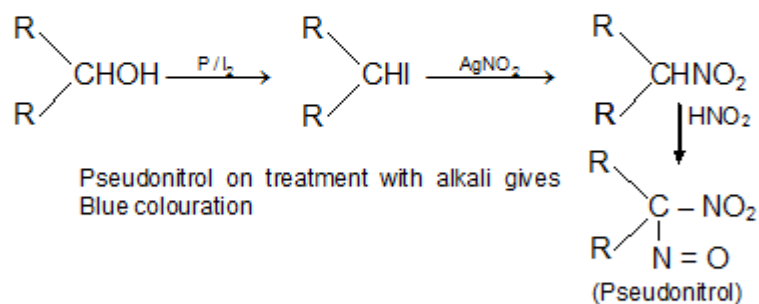
3° Alcohol: R<sub>3</sub>CHOH + ZnCl<sub>2</sub> + HCl → R<sub>3</sub>CHCl white turbidity instantaneously.

## 2. Victor Meyer Test

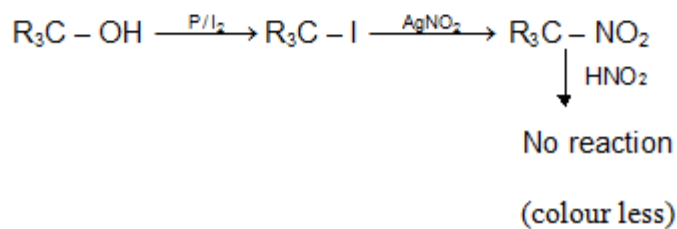


Nitrolic acid on treatment with alkali gives colouration

2° alcohol:



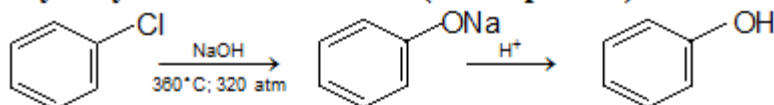
3° alcohol:



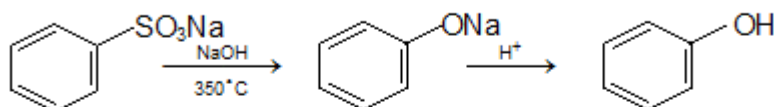
Phenols:

### Preparation:

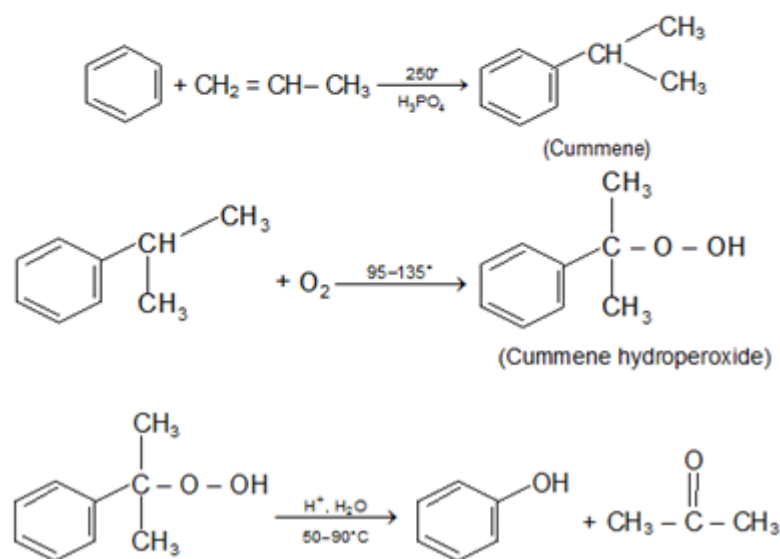
#### a. Hydrolysis of chlorobenzene: (Dow's process)



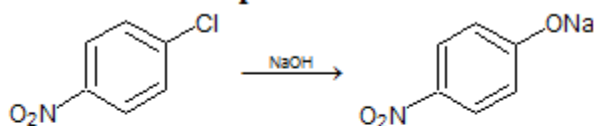
#### b. Alkali fusion of Sodium benzene sulfonate



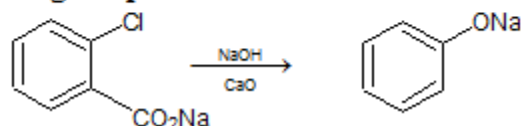
#### c. From Cumene Hydroperoxide



#### d. Aromatic Nucleophilic Substitution of Nitro Aryl Halides



#### e. Distillation of phenolic acids with soda-lime produces phenols, e.g. sodium salicylate gives phenol.



### Physical Properties of Phenols

Phenol is a colorless, toxic, corrosive, needle shaped solid.

Phenol soon liquifies due to high hygroscopic nature.

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Phenol is less soluble in water, but readily soluble in organic solvents.

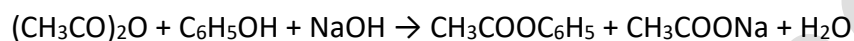
Simplest phenols, because of hydrogen bonding have quite high boiling points.

o-nitrophenol is, steam volatile and also is less soluble in water because of intramolecular hydrogen bonding

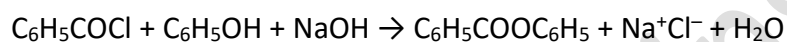
## Chemical Properties of Phenols

### a) Formation of Esters

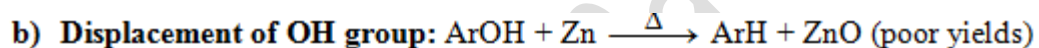
Phenyl esters (RCOOAr) are not formed directly from RCOOH. Instead, acid chlorides or anhydrides are reacted with ArOH in the presence of strong base



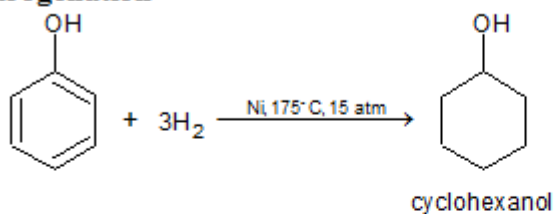
Phenyl acetate



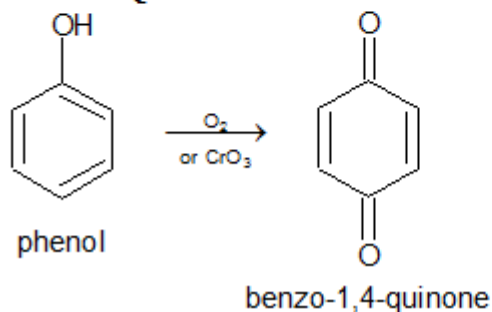
Phenyl benzoate



### c) Hydrogenation

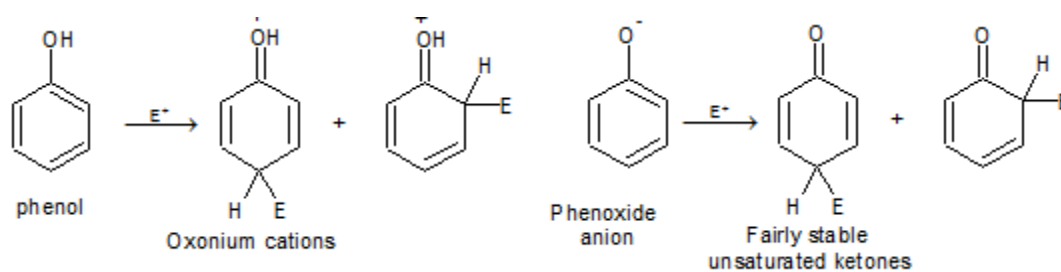


### d) Oxidation to Quinones



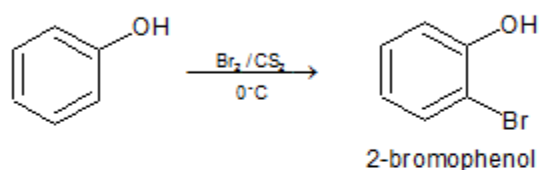
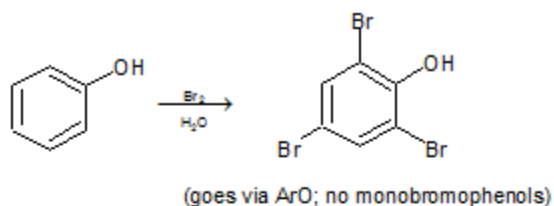
**e) Electrophilic Substitution** The —OH and even more so the —O(phenoxide) are strongly activating ortho, para - directing

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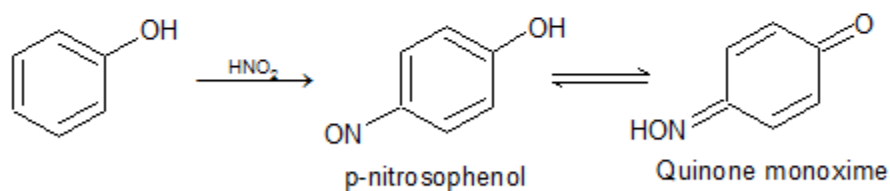


Special mild conditions are needed to achieve electrophilic monosubstitution in phenols because their high reactivity favors both polysubstitution and oxidation

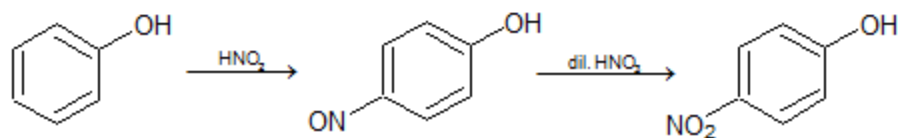
### f) Halogenation



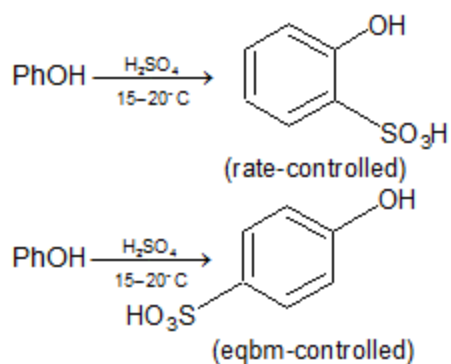
### h) Nitrosation



### i) Nitration

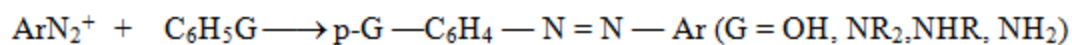


**j) Sulfonation**

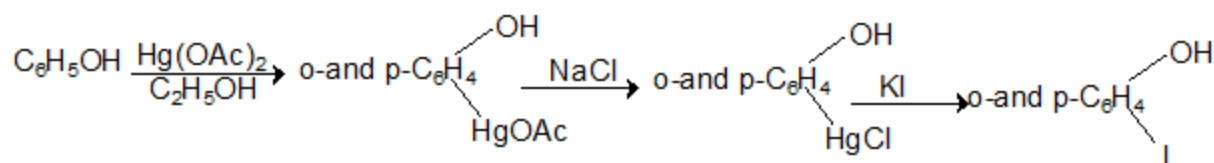


**k) Diazonium salt coupling to form azophenols**

**Coupling** (G in ArG is an electron – releasing group)



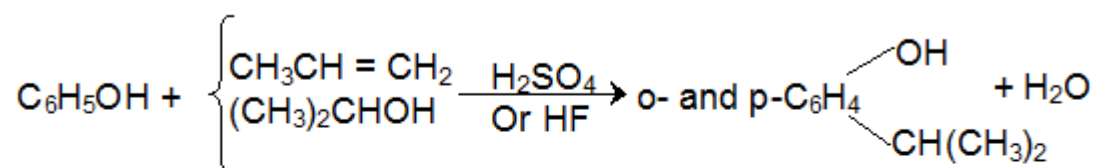
**l) Mercuration**



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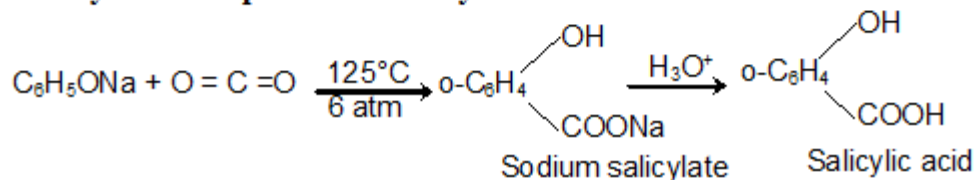


**m) Ring alkylation**



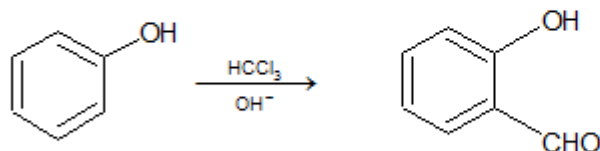
RX and  $\text{AlCl}_3$  give poor yields because  $\text{AlCl}_3$  coordinates with O.

**n) Kolbe synthesis of phenolic carboxylic acids**



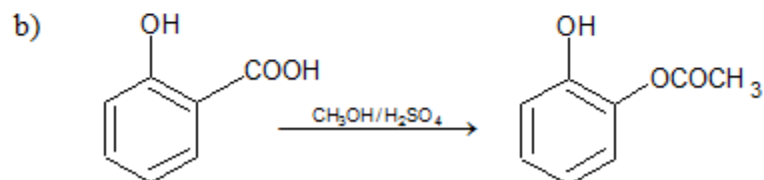
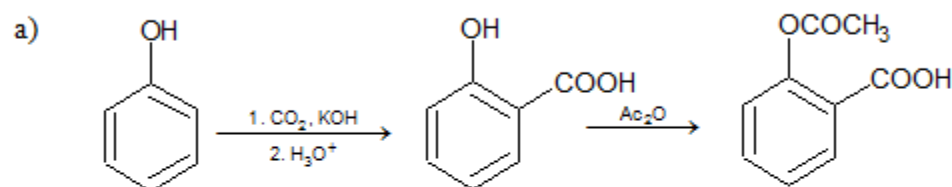
Phenoxide carbanion adds at the electrophilic carbon of  $\text{CO}_2$ , para product is also possible.

**o) Reimer – Tiemann synthesis of phenolic aldehydes**



The electrophile is the dichlorocarbene,  $\text{:CCl}_2$ , formation of carbene is an example of  $\alpha$ -elimination.  $\text{OH}^- + \text{HCCl}_3 \xrightarrow{-\text{HCl}} \text{:CCl}_2$

**p) Synthesis of (a) aspirin (acetylsalicylic acid) (b) oil of wintergreen (methyl salicylate)**

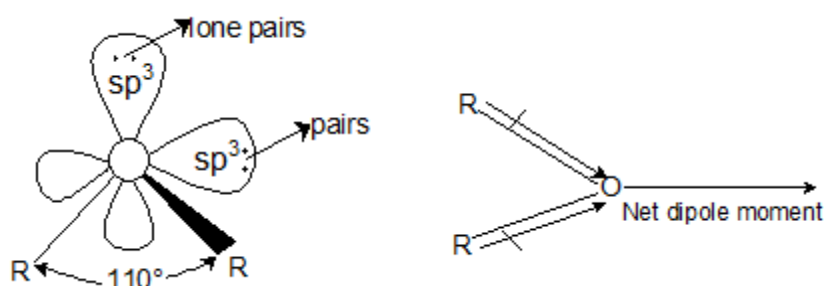


## Ethers:

### Physical Properties of Ethers

**Physical state, colour and odour:** Dimethyl ether and ethyl methyl ether is gas at ordinary temperature while the other lower homologues of ethers are colourless liquid with characteristic 'ether smell'.

**Dipole nature:** Ethers have a tetrahedral geometry i.e., oxygen is  $sp^3$  hybridized. The C—O—C angle in ethers is  $110^\circ$ . Because of the greater electronegativity of oxygen than carbon, the C—O bonds are slightly polar and are inclined to each other at an angle of  $110^\circ$ , resulting in a net dipole moment.

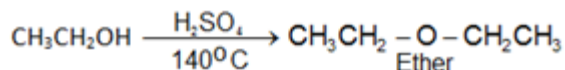


Bond angle of ether is greater than that of tetrahedral bond angle of  $109^\circ 28'$ .

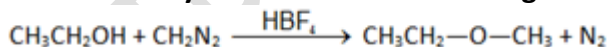
**Solubility and boiling point:** Due to the formation of less degree of hydrogen bonding, ethers have lower boiling point than their corresponding isomeric alcohols and are slightly soluble in water.

### Preparation of Ethers:

#### a) From alcohols:



Order of dehydration of alcohol leading to formation of ethers:  $1^\circ > 2^\circ > 3^\circ$

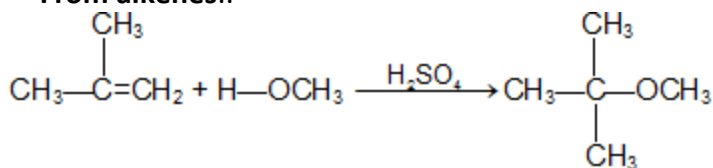


#### b) Williamson's synthesis:

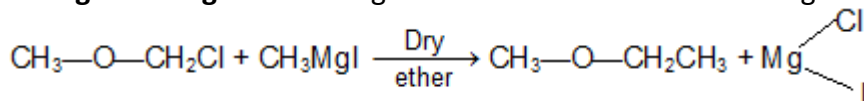


In case of tertiary substrate elimination occurs giving alkenes.

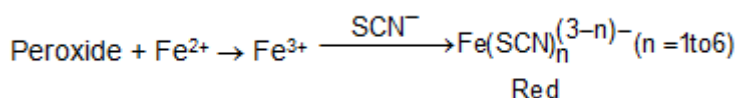
From alkenes:



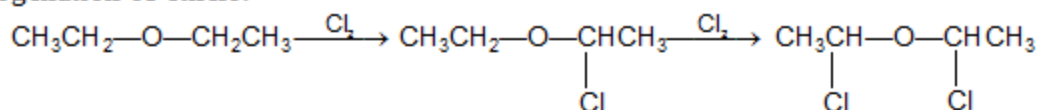
From Grignard reagent: Treating a - halo ethers with suitable Grignard reagents.



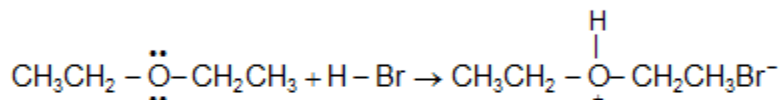
On standing in contact with air, most aliphatic ethers are converted slowly into unstable peroxides. The presence of peroxides is indicated by formation of a red colour when the ether is shaken with an aqueous solution of ferrous ammonium sulfate and potassium thiocyanate



f) Halogenation of ethers:



g) Ethers as base:

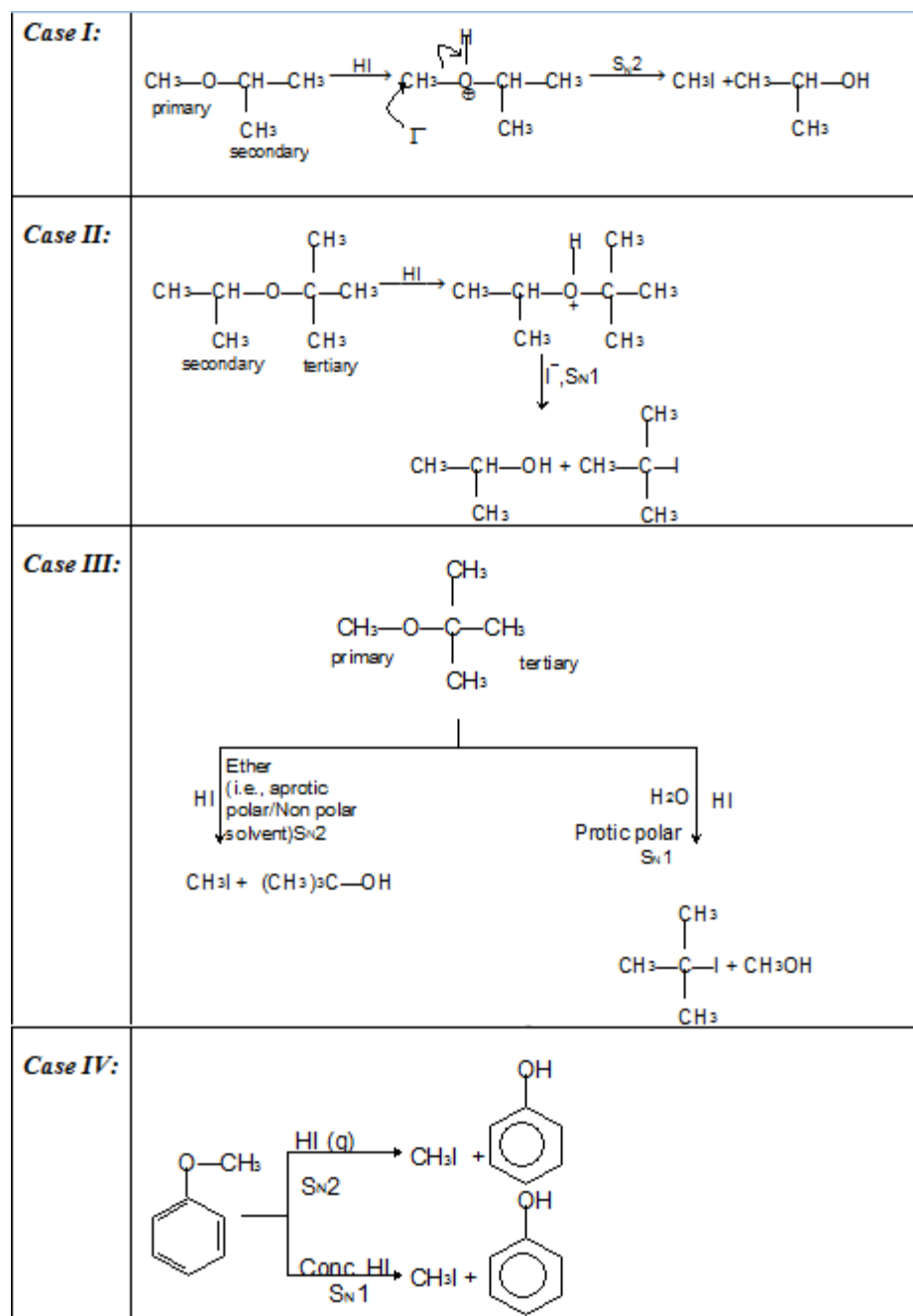


h) Reaction With Cold conc. HI/HBr:

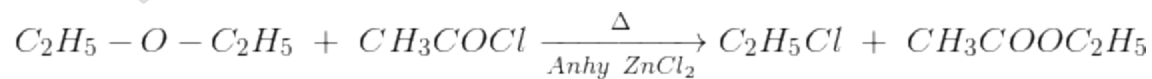


i) Hot conc. HI/HBr:



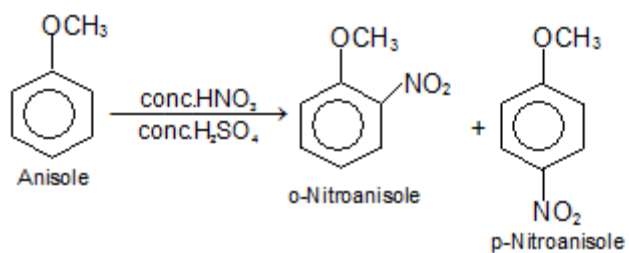
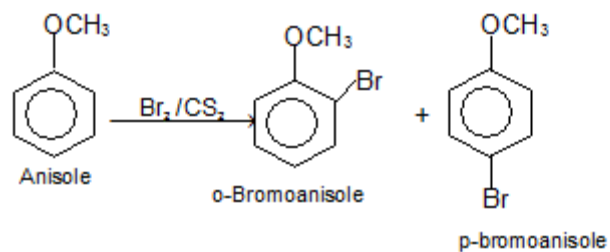


**i) Reaction with acid chlorides and anhydrides:**



**j) Electrophilic substitution reactions**

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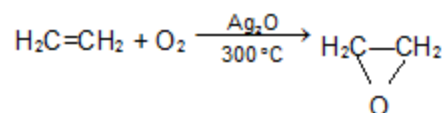


**Epoxides or Oxiranes:**

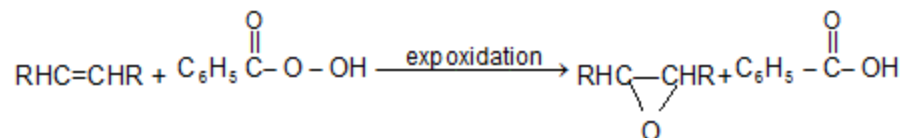
**Preparation**

a) Oxidation of ethylene :

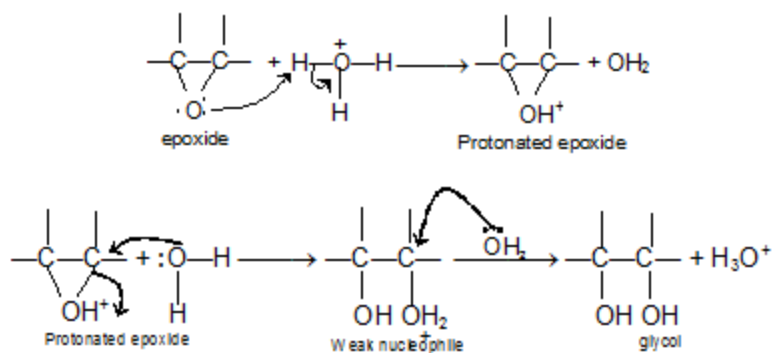
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b) Exoxidation :



Acid catalysed ring opening



Base catalysed ring opening:

